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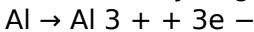
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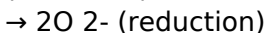
Oxidation

Reduction

electrons or the
addition of oxygen;
reduction is the gain of
electrons or the
addition of hydrogen.



(oxidation); $\text{O}_2 + 2\text{e}^{-}$



(answers will vary)

5.5: Oxidation- Reduction (Redox) Reactions - Chemistry ...

Answer. Reduction: $\text{Ca}^{2+} + 2\text{e}^{-} \rightarrow \text{Ca}$.

Oxidation: $2(\text{K} \rightarrow \text{K}^{+} + \text{e}^{-})$

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Oxidation

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+ e⁻) Combined: Ca

$2+ + 2K \rightarrow Ca + 2K +$

The substance oxidized is the reactant that had undergone oxidation:

K; The substance reduced is the reactant that had undergone

reduction: Ca²⁺ The reducing agent is the

same as the substance oxidized: K

13.1: Oxidation-Reduction (Redox) Reactions - Chemistry ...

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Questions pertaining to redox reactions. If you're seeing this message, it means we're having trouble loading external resources on our website.

Redox reactions questions (practice)
| Khan Academy

reduction: a process that involves a complete or partial gain of electrons or the loss of oxygen; it

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results in a decrease in the oxidation number of an atom: oxidation number: a positive or negative number assigned to a combined atom according to a set of arbitrary rules: oxidation: a process that involves complete or partial loss of electrons or a gain of oxygen; it results in an increase in the oxidation number of an atom: redox reaction:

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another name for an
oxidation-reduction
reaction

Answers

**Quia - Chapter 20
"Oxidation-
Reduction
Reactions"**

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its home. explain why
the mass of a rusted
nail would

Oxidation And Reduction Practice With Answers

The sum of the
oxidation numbers of
the elements in a
neutral compound is 0 .
In a polyatomic ion,
however, the sum is

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Oxidation

Reduction

equal to the Oxidation
'numbers help you
keep track of 6 transfer
in redox reactions. An
oxidation number
increase is 7 , while a 8
is reduction.

honors-redox copy - Mister Chemistry

Definition of oxidation:
species loses electrons
(electrons are a
product) Definition of
reduction: species
gains electrons
(electrons are a

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Oxidation

Reduction

reactant) Determine whether the reaction is an oxidation or a reduction of iron .

Electrons are a reactant. Fe^{3+} gains 3 electrons to produce $\text{Fe}(s)$ Therefore this is a reduction reaction.

Oxidation and Reduction Concepts Chemistry Tutorial

Reduction is when the total number of electrons increases in a reaction; oxidation is

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Reduction

when the total number of electrons decreases in a reaction.

Reduction is a reaction that removes an electron ...

**Redox Reactions -
Practice Test
Questions & Chapter
Exam ...**

Oxidation and reduction are two types of chemical reactions that often work together.

Oxidation and

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Oxidation

Reduction

reduction reactions involve an exchange of electrons between reactants. For many students, the confusion occurs when attempting to identify which reactant was oxidized and which reactant was reduced.

What is the Difference Between Oxidation and Reduction?

Oxidation numbers are very important and are

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Concept Review

Answers

used for 1) naming compounds, 2) balancing oxidation-reduction reactions, 3) calculations in electrochemistry and other areas of chemistry. Rule 0 The following rules are in the form of a hierarchy; that is, the first stated rule takes precedence over subsequent rules if a conflict arises.

Oxidation Number

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Oxidation

Reduction

Exercise - Multidict

Method 1: Oxidation

number method 1.

Assign oxidation

numbers to all

elements in the

reaction 2. From the

changes in O.N.,

identify the oxidized

and reduced species 3.

Compute the number

of electrons lost in the

oxidation and gained in

the reduction from the

O.N. changes 4.

Multiply one or both of

these numbers by

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appropriate

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Academic Resource

Center

Oxidation-state change
Comprehensive
definitions of oxidation
and reduction have
been made possible by
modern molecular
structure theory. Every
atom consists of a
positive nucleus,
surrounded by
negative electrons,
which determine the
bonding characteristics

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Oxidation

Reduction

of each element. In forming chemical bonds, atoms donate, acquire, or share electrons.

Oxidation-reduction reaction - General theory | Britannica

The concept of reduction has undergone a similar evolution. At high temperature, zinc oxide, ZnO , reacts with carbon, C , to form molten zinc and carbon

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Oxidation

Reduction

monoxide gas. $\text{ZnO(s)} + \text{C(s)} \rightarrow \text{Zn(l)} + \text{CO(g)}$

Bonds between zinc atoms and oxygen atoms are lost in this reaction, so chemists say the zinc has been reduced.

Chapter 9 - An Introduction to Chemistry: Oxidation ...

Oxidation-reduction reactions are of central importance in organic chemistry and

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Oxidation

Reduction

biochemistry. The burning of fuels that provides the energy to maintain our civilization and the metabolism of foods that furnish the energy that keeps us alive both involve redox reactions. All combustion reactions are also redox reactions.

**Redox Reactions in
Organic Chemistry
and Biochemistry**

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REDOX: Oxidation and Reduction Interactive Notebook Students will be actively involved in learning about Oxidation and Reduction by creating this interactive notebook! There are 4 lessons included, each with their own shapes, to cover the new concepts learned. Students will be cutting and pasting t

Oxidation And

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Worksheets &

Teaching Resources

| TpT

Most oxidation-reduction (redox) processes involve the transfer of oxygen atoms, hydrogen atoms, or electrons, with all three processes sharing two important characteristics: (1) they are coupled—i.e., in any oxidation reaction a reciprocal

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Oxidation

Reduction

reduction occurs, and (2) they involve a characteristic net chemical change—i.e., an atom or electron goes from one unit of matter to another.

Oxidation-reduction reaction | chemical reaction | Britannica

Review of reduction and oxidation opens this video. It really helps stimulate understanding of these biochemistry topics. By

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Answers
explaining the loss or gain of electrons, Salman Khan is able to clarify the concept of respiration processes.

Oxidation and Reduction Lesson Plans & Worksheets

Chemical reactions in which electrons are transferred are called oxidation-reduction, or redox, reactions.

Oxidation is the loss of electrons. Reduction is the gain of electrons.

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Oxidation

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Oxidation and reduction always occur together, even though they can be written as separate chemical equations.

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