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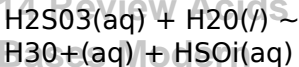
Chapter 14 Review Acids Bases

CHAPTER 14 REVIEW .
Acids and Bases.

SHORT ANSWER

Answer the following
questions in the space
provided. 1. a. Write
the two equations that
show the two-stage
ionization of sulfurous
acid in water. stage 1:

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b. Which stage of ionization usually produces more ions? Explain your answer.

CHAPTER 14 REVIEW Acids and Bases - Weebly

CHAPTER FOURTEEN
ACIDS AND BASES
CHAPTER 14 ACIDS
AND BASES 5 When
 H_3PO_4 is added to
water, the three acids
that are present are

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Chemistry

H_3PO_4 , H_2PO_4^- , and
 HPO_4^{2-} — H_3PO_4 , with
the largest K_a value, is
the strongest of these
weak acids The
conjugate

[DOC] Chapter 14 Review Acids And Bases Mixed

CHAPTER 14 REVIEW
Acids and Bases

SECTION 1 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

Name the following

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compounds as acids:
sulfuric acid a. H_2SO_4
sulfurous acid b. H_2SO_3
3 hydrosulfuric acid c.
 H_2S perchloric acid d.
 HClO_4 hydrocyanic
acid e. hydrogen
cyanide 2. H_2S Which
(if any) of the acids
mentioned in item 1
are binary acids? 3.

14 Acids and Bases

-Acids are defined as
any substance that can
accept an electron pair
to form a covalent

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bond (Substances do NOT have to contain Hydrogen) -Bases are defined as any substance that donates an electron pair

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CHAPTER 14 REVIEW.

Acids and Bases.

SHORT ANSWER

Answer the following questions in the space provided. 1. Answer the following questions

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according to the
Brønsted-Lowry
definitions of acids and
bases: HSO_3^- . a.
What is the conjugate
base of H_2SO_4 .

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to Review #7: Acids,
Bases and Salts

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14 Review Acids Answers to Review #7:

Acids, Bases and Salts

1. Know the meanings
of and be able to apply
the following terms:

Bronsted-Lowry acid

strong acid

concentrated chemical

...

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chapter of this Holt
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Review Questions -

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work step by step

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Learning

Chapter 14 - Acids and Bases - Review

Page 12/24

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**Questions - Page
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When an acid and a base are mixed, the H^+ from the acid combines with the OH^- from the base to form H_2O . Acid-Base Titration In a titration, a substance in a solution of KNOWN concentration is reacted with another substance in a solution of UNKNOWN concentration.

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Acids and Bases

Chapter 14
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CHAPTER 14 ACIDS
AND BASES 5 When
 H_3PO_4 is added to
water, the three acids
that are present are
 H_3PO_4 , H_2PO_4^- , and
 HPO_4^{2-} . H_3PO_4 , with
the largest K_a value, is
the strongest of these
weak acids. The
conjugate bases of the
three acids are H_2PO_4^- ,
 HPO_4^{2-}

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**CHAPTER FOURTEEN
ACIDS AND BASES**

Proton transfer
reactions favor the
production of
[stronger/weaker]
acids and bases

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and Bases
Flashcards**

CHAPTER FOURTEEN
ACIDS AND BASES For
Review 1. ... CHAPTER
14 ACIDS AND BASES 5
When H_3PO_4 is added
to water, the three

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acids that are present are H_3PO_4 , H_2PO_4^- , and HPO_4^{2-} . H_3PO_4 , with the largest K_a value, is the strongest of these weak acids. The conjugate ...

Modern Chemistry Chapter 14 Review Answers Acids And Bases

Acids and Bases:

Chapter 14 & 15 . HW:

•Read Ch 14: •Fill in as much of the acid base table as you can, as

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you read . Acid base
conductivity and
reactivity Conductivity,
Reac'tivity,, ... 3. Work
on ch 14-15 review
(due Wednesday, test
Friday!) • Solutions in
which [H

Acids and Bases: Chapter 14 & 15

Chapter 14: Acids and
Bases. Ch 14 Page 1.
Arrhenius Base -a
hydroxide containing
compound that
dissociates to produce

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OH-when dissolved in
water $\text{NaOH}(\text{aq})$
 $\rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$
Bases that contain OH-
are ionic compounds.
Ionic substances
dissociate in water.

Chapter 14: Acids and Bases

pH of a weak base: pH
of .2 M of NH_3 (weak
base). Conjugate acids
and bases :

Introduction to
conjugate acids and
bases pK_a and pK_b

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relationship : The pK_a
and pK_b relationship
between conjugate
acids and bases (both
of which are weak).

Chapter Fourteen [Acids and Bases] - Wattsburg

Modern Chemistry 121

Acids and Bases

CHAPTER 14 REVIEW

Acids and Bases

SECTION 3 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

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Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: _____ a. What is the conjugate base of H_2SO_3 ? _____ b.

CHAPTER 14 REVIEW Acids and Bases

equivalent amounts of acids and bases react, the three properties just described disappear because the acid is “neutralized.” The

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reaction products are water and an ionic compound called a salt. 5. Acids conduct electric current. ... 3. CHAPTER 14 4. Bases + ...

CHAPTER 14 Acids and Bases

Chapter 14 Review:
Acids and Bases.

Section 14.1 Bronsted Acids and Bases. Acids - molecules that can lose H^+ (proton donors) making H_3O^+ in water

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(acids lose H^+) Bases - molecules than can gain H^+ (proton acceptors) and/or make OH^- in water (bases gain H^+ / make OH^-) Conjugate acid/base pairs - differ by one H^+ .

Chapter 15 Review Acid Base Equilibria

Chapter 14 Acids Bases
Answers Chapter 14
Acids Bases Answers
As recognized,
adventure as capably

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