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To determine a  
compound/element's

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molar mass: 1) Make a list of all the elements and the number of that specific element. 2) Use the periodic table to write down the mass for each element. 3) After you have determined the mass and number of each element, you will multiply the mass by the number of atoms. 4) Once you have multiplied, you must add all the numbers together to get your

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Quantities Section 10.1

- The Mole: A

Measurement of Matter

You often measure the  
amount of something

by count, by mass, or

by volume. A mole

(mol) of a substance is

$6.02 \times$

$10^{23}$  representative

particles of that

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substance.

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answers. 5.

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Quantities BASICS: •

The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole



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was founded by a  
scientist named  
Avagadro, and he  
decided to use the

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THE MOLE AND

QUANTIFYING MATTER

10.1 The Mole: A ...

After reading Lesson

10.1, answer the

following questions. ...

The quantities 1 mol

and 22.4 L can be used

in conversion factors

that change moles to

volume and volume to

moles at STP.

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Answers 10 1 kg 103 g

368 Chapter 10

Chemical Calculations  
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Equations. 10.1

Equation Stoichiometry

369 The ratio of moles

of P 4O 10 to moles of

P (which came from

the subscripts in the

chemical formula, P 4O

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