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Investigation 11A: Stoichiometry Part 1:
The reaction. 4 Investigation 11A:
Stoichiometry 4. Slowly pour the vinegar
into the baking soda. Pouring it too fast
might cause the bubbling to overflow!
Wait until the bubbling has stopped and

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record the mass. ... Answers in as fast as
15 minutes ...

Inv11A_Slides.ppt - Investigation 11A Stoichiometry If ...

Stoichiometry is the tool for answering
these questions. Stoichiometry The
study of quantitative relationships
between the amounts of reactants used

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and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

Chapter 11: Stoichiometry

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X 101 grams . 2. 1.151 grams 4. 13
moles 6. 8.4 X 10²² atoms 8. 1.80
grams DATA TABLE: mass of empty
evaporating dish and watch glass mass
...

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Investigation 11A: THUNDERSTORMS 1.
weak strong 2. falls rises 3. cold warm 4.
high low Applications 5. 78 75 67 6. 26
31 48 7. high low 8. northward

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southward 9. southerly northerly 10.
weak strong 11. widespread light rainfall
local strong thunderstorms 12. were
were not 13. were were not 14.

lab11 - Investigation 11A **THUNDERSTORMS 1 weak strong ...**

Stoichiometry problems can be
characterized by two things: (1) the

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information given in the problem, and (2) the information that is to be solved for, referred to as the unknown . The given and the unknown may both be reactants, both be products, or one may be a reactant while the other is a product.

Stoichiometry | Chemistry for Non-

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Majors

Grade 11A Science Related
Reading/Chemistry 1 Stoichiometry and
Safety Air Bags. UTask (1): Pre -reading
activity: UChoose the correct answer: 1-
which law describes the relation
between the temperature and the
pressure of a gas: a- Charles's law . b-
Boyle's law . c- Gay-lussac's law . 2-how

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increasing temperature affects the gas's pressure?

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ANSWERS: 8A. NH_4NO_3 8B. 0.125 mol
 $(\text{NH}_4)_3\text{PO}_4$, 0.375 mol NaNO_3 8C. 29.5 g
 Na_3PO_4 9A. FePO_4 9B. 0.149 mol
 $\text{Ca}_3(\text{PO}_4)_2$, 0.149 mol $\text{Fe}_2(\text{CO}_3)_3$ 9C.

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55.2 g CaCO_3 10. What quantity of energy is released when $2 \text{C}_5\text{H}_{10} + 15 \text{O}_2 \rightarrow 10 \text{CO}_2 + 10 \text{H}_2\text{O} + 6,199 \text{ kJ}$. 540 g of cyclopentane are burned? 11A. What mass of iron is needed to react with 16.0 grams of sulfur?

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Stoichiometry Answers

Stoichiometry is the quantitative balancing of elements in chemical reactions. Conservation of mass requires that all atoms that enter a reaction as reactants must exit the reaction in the products. The Ideal Gas Law is used to model equilibrium conditions of most gases, relating the pressure, volume,

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temperature, and moles of gas.

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